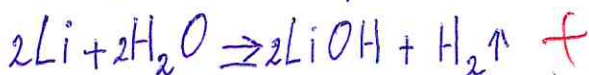
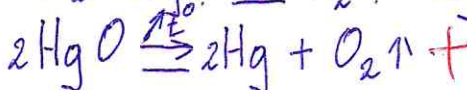
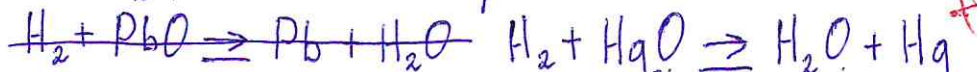


№ 9. 1. А - Li, т.к. он самый легкий из всех металлов.



В - H_2 , т.к. он легко горит. -



С - Hg, т.к. ртуть - серебристо-серая жидкость.

Ответ: А - Li; В - H_2 ; С - Hg.

1 45 М. Д.
2 45 М. Д.
3 35 М. Д.
4 35 М. Д.
5 55 М. Д.

№ 9. 2. Дано:

Кол-во атомов - 220 шт.

$m(\text{CaCO}_3) = 52 \text{ г/шт.}$

Найти:

$m(\text{Ca}) = ?$

Решение

$$J = \frac{m}{M} \quad m = J \cdot M$$

$$M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 40 + 12 + 48 = 100 \text{ г/моль}$$

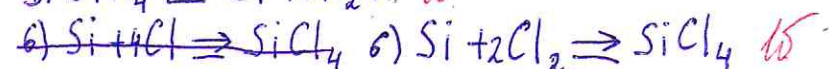
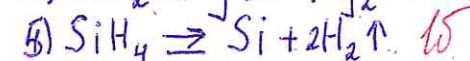
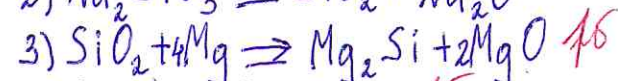
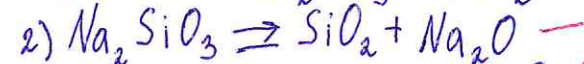
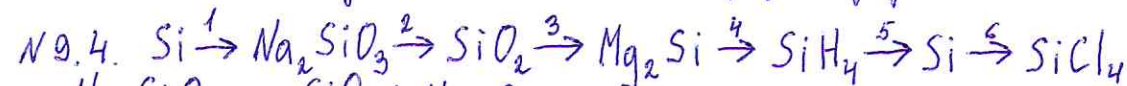
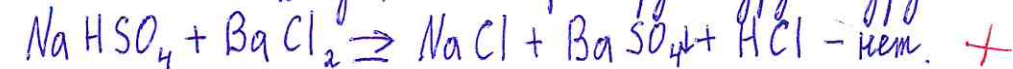
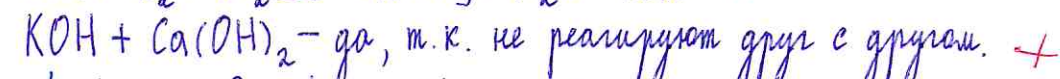
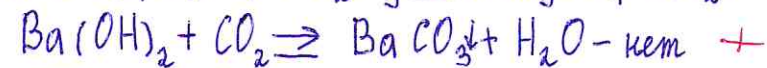
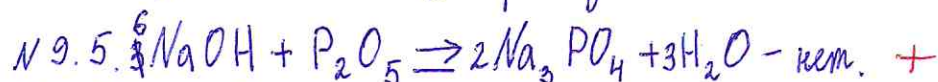
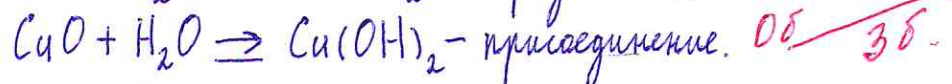
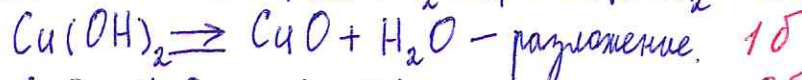
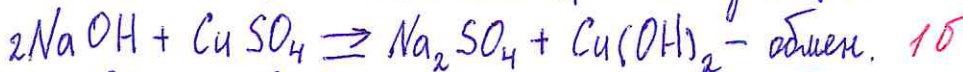
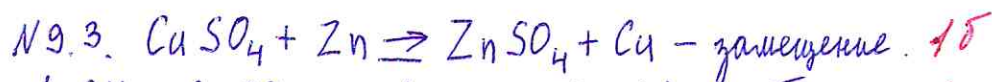
$$J(\text{CaCO}_3) = m(\text{CaCO}_3) = 220 \cdot 5 = 1100 \text{ г}$$

$$J(\text{CaCO}_3) = \frac{1100}{100} = 11 \text{ моль}$$

$$m(\text{Ca}) = 40 \cdot 11 = 440 \text{ г}$$

Ответ: 440 г.

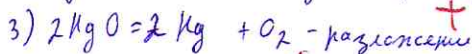
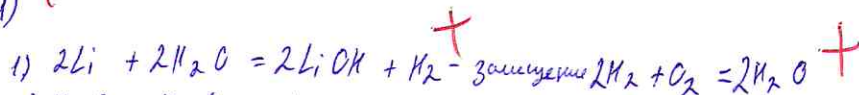
45



35

1 55 М
2 45
3 25
4 - 3
5 15
ШИФР 9-22

1) 15



Ответ: А - Li

В - H_2

С - Hg

2. Дано:

$$m(\text{CaCO}_3) = 52$$

(Ca) - ?

Решение:

$$M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 100 \text{ г/мол}$$

$$w(\text{Ca}) = \frac{40}{100} = 0,4$$

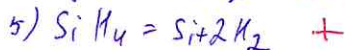
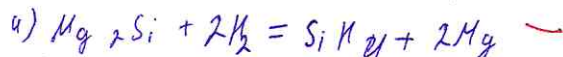
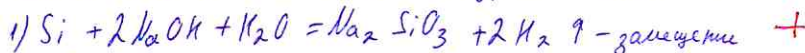
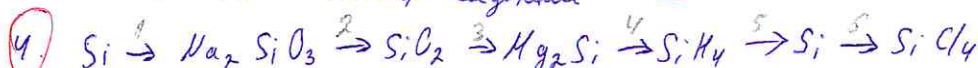
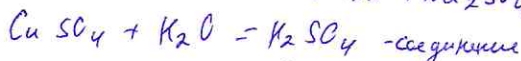
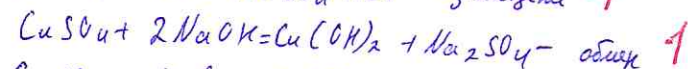
$$m(\text{Ca}) = w(\text{Ca}) \cdot m(\text{CaCO}_3) = 0,4 \cdot 52 = 20,8$$

$$20,8 \cdot 20 = 416$$

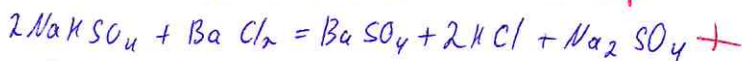
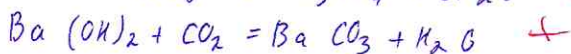
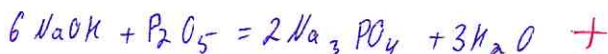
Ответ: 416

45

3. NaOH ; CuSO_4 ; Zn



5.



46

Бланк ответа

Класс 9Б⁴

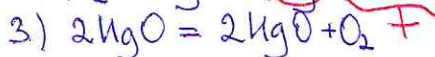
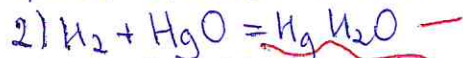
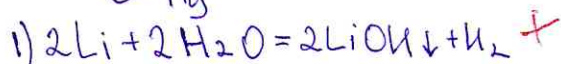
ШИФР 9-66

1 38
2 45
3 25
4 - 5
5 18

9.1. А - Li

В - H₂

С - Hg



9.2. Dano:

За рог = 220 еми

 $m(\text{CaCO}_3)_{\text{в рог}} = 52$

(Ca) в рог — ?

Решение: $M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 100$ $M(\text{Ca}) = 40$

100 : 240 =

52 : x

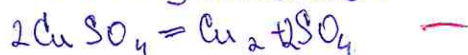
$$x = \frac{5 \cdot 40}{100} = 22$$

220 : 2 = 440 — за рог восстановления

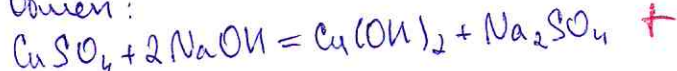
Омбем: 4402.

45

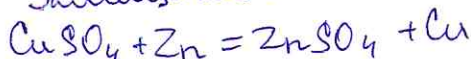
9.3. Разложение:



Омбем:

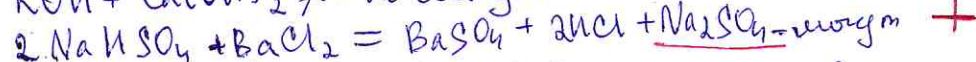
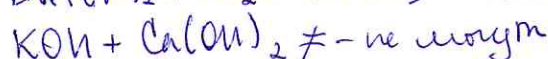
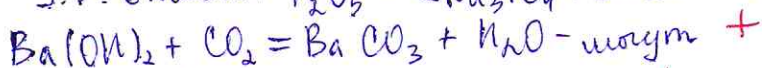
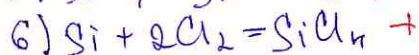
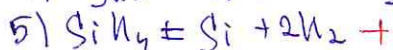
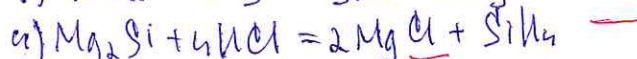
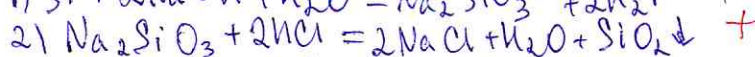
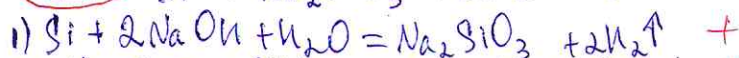
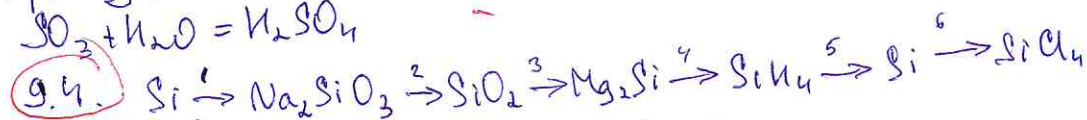


Замещение:

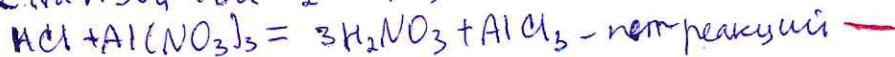


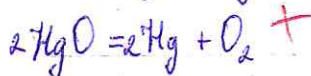
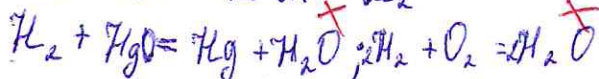
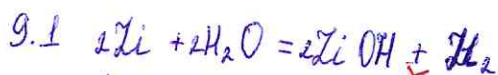
25

Амурежение:



48





$$9.2 \quad M_r(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 100$$

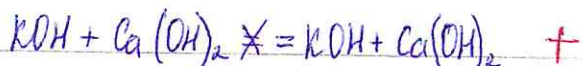
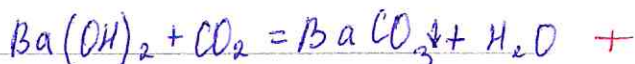
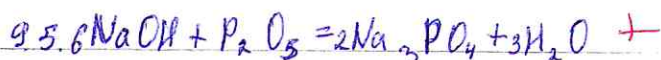
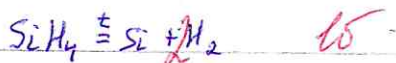
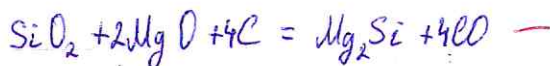
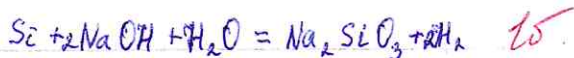
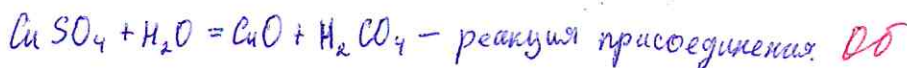
В 100г (CaCO_3) содержится 40г Ca .

В 5г (CaCO_3) содержится $x\text{г}$ Ca

$$2) \quad x = \frac{5 \cdot 40}{100} = \frac{40}{20} = 2\text{г}$$

$$3) \quad 220\text{г} \cdot 2\text{г} = 440\text{г}$$

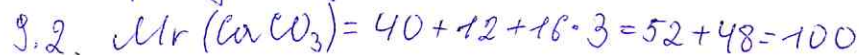
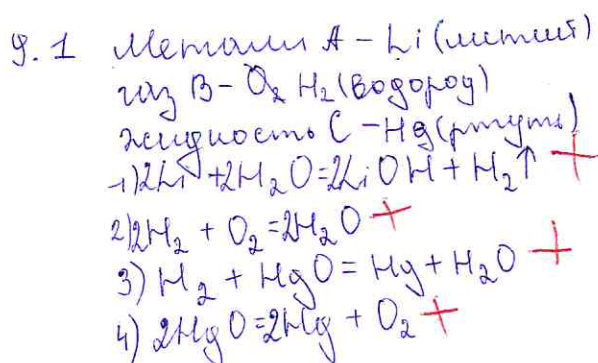
Ответ: 440г за год. (CaCO_3)



$\text{HCl} + \text{Al(NO}_3)_3 \nrightarrow$ могут если они не вступают в химическую реакцию. $+$

58

1 58 M
2 45 H
3 25 H
4 -25 H
5 58 H

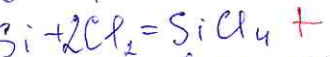
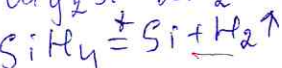
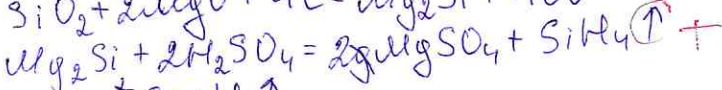
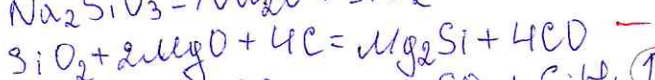
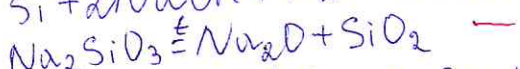
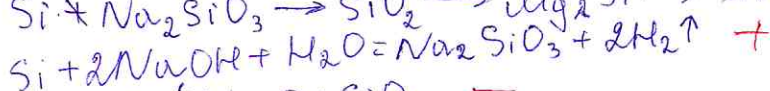
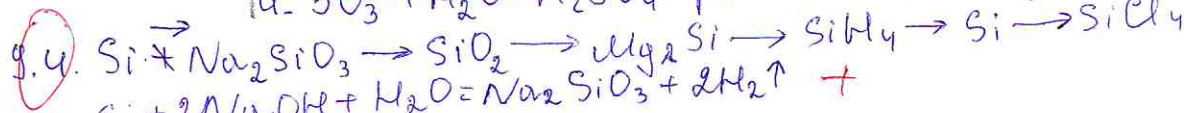
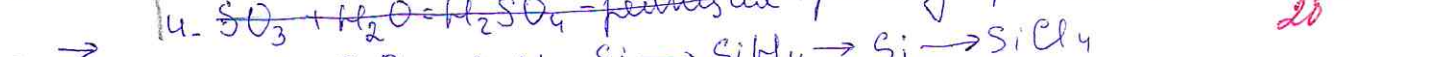
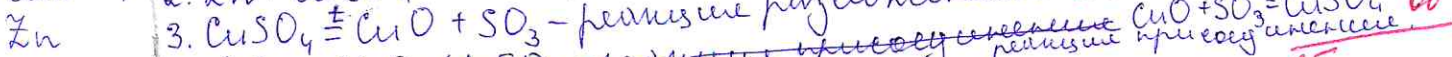
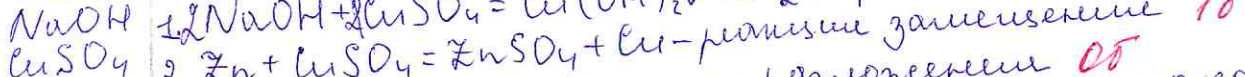
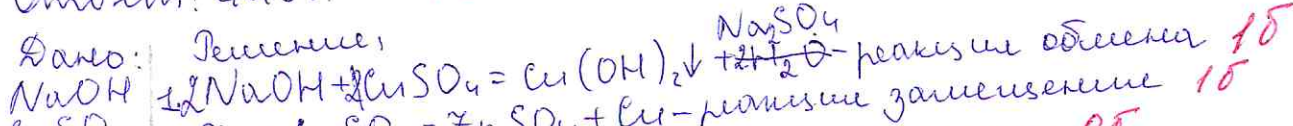


1) $\frac{5 \cdot 40}{100} = \frac{200}{100} = 2г$

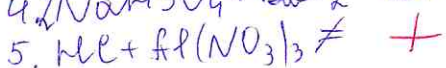
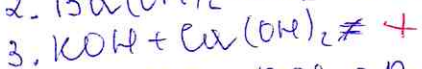
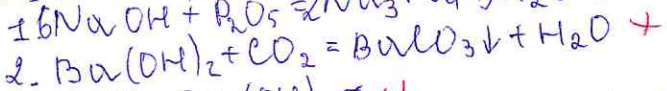
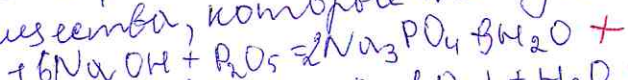
2) $2 \cdot 220 = 440г$

Ответ: 440г.

9.3. Дано: Решение

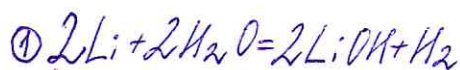


9.5. В растворе одновременно могут находиться только те вещества, которые не взаимодействуют между собой.

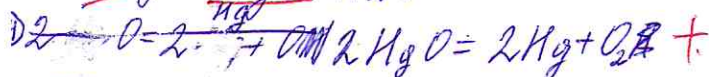


56

1. 55
 2. 45
 3. 25
 4. -2
 5. 55

9.1.

А-литий (Li)

В-водород (H₂)С- ~~литий~~ (Mg) (H₂)9.2.

$$M(\text{CaCO}_3) = \frac{110 + 12 + 48}{40 + 12 + 48} = 100$$

Сколько Ca содержится в карбонате кальция (CaCO₃)

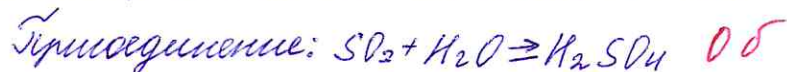
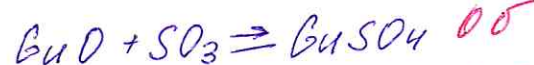
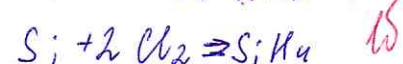
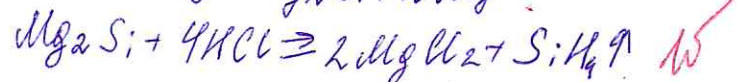
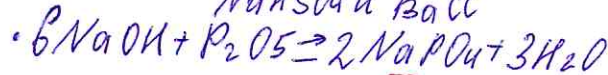
$$\text{в } 100\text{г.} = 40\text{г. (Ca)}$$

$$\text{в } 5\text{г.} = x$$

$$x = 5 \cdot 4 : 100 = 2\text{г.}$$

2 · 220 = 440 (г) - Ca необходимо выплавить за 1 год.

Ответ: 440 г

9.3.9.4.9.5.Могут только две пары (не взаимодействуют между собой)
KOH + NaOH9.5.Могут: KOH и Ca(OH)₂
HCl и H(NO₃)₃Не могут: NaOH и P₂O₅Ba(OH)₂ и CO₂NaHSO₄ и BaCl₂

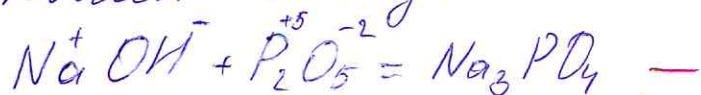
(65)

15

1 45
2 45
3 25
4 65
5 15

9.5.

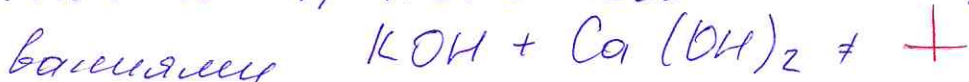
1. может, т.к. основания реагируют с кислотными оксидами



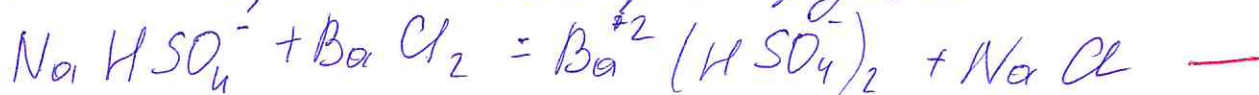
2. не может, т.к. основания не реагируют с оксидами



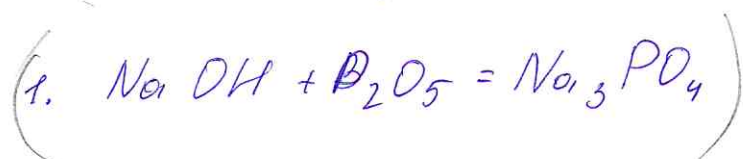
3. не может, т.к. основания не реагируют с основаниями



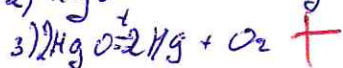
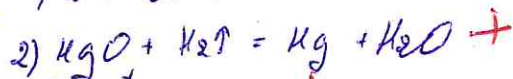
4. может, т.к. соли реагируют с солями



5. не может, т.к. всё растворяется



18



Ответ: А - Li, В - H₂↑; С - Hg.

2. Дано: 220 амс.
 $\frac{m(\text{CaCO}_3) = 52}{m(\text{Ca}) = ?}$

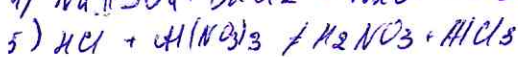
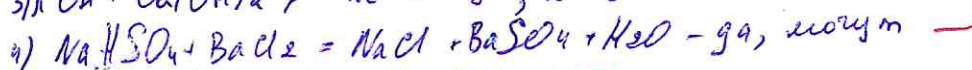
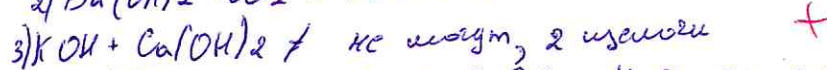
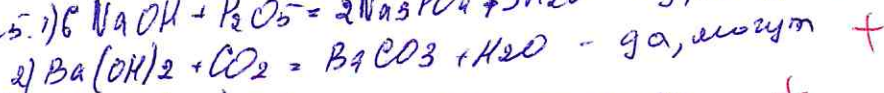
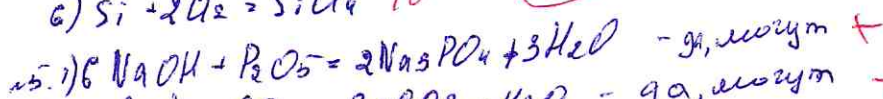
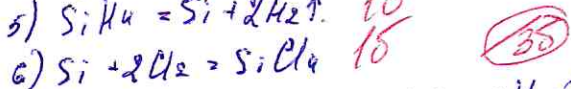
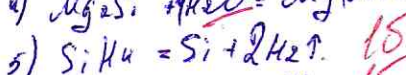
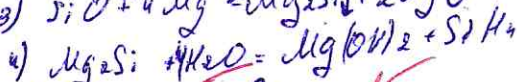
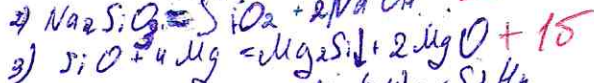
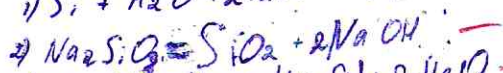
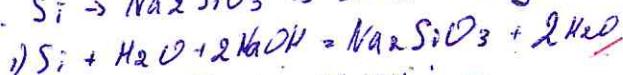
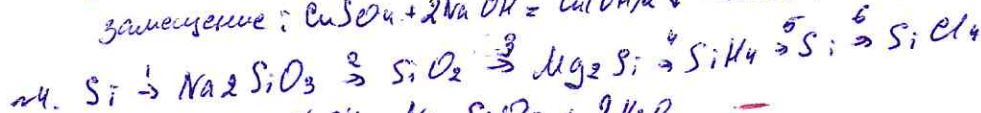
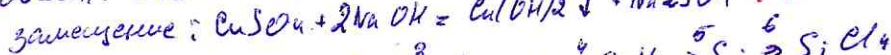
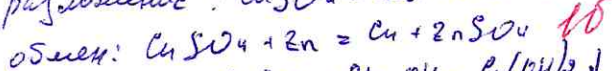
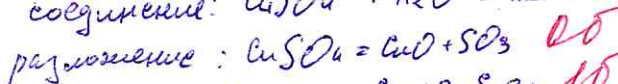
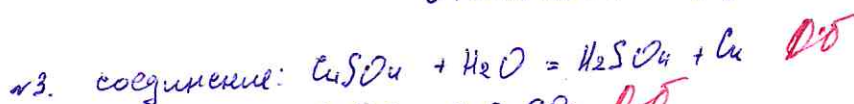
$M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 100 \text{ г/моль}$

$w(\text{Ca}) = \frac{m(\text{Ca})}{M} = \frac{40}{100} = 0,4$

$m(\text{Ca}) = m(\text{CaCO}_3) \cdot w(\text{Ca}) = 0,4 \cdot 52 = 20,8$

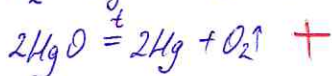
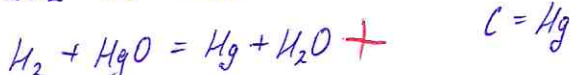
$220 \cdot 2 = 440 \text{ г.}$

Ответ: 440 г. - m(Ca)



1. 45
 2. 40
 3. 25
 4. - 30
 5. 35

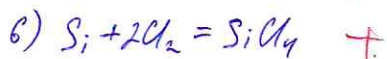
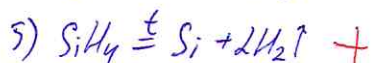
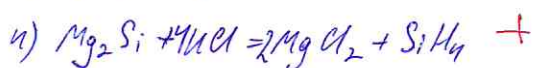
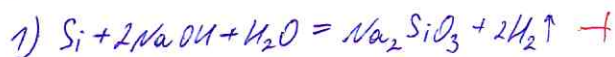
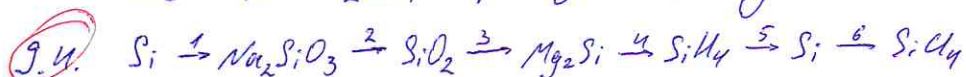
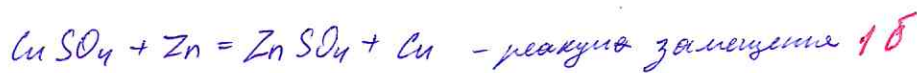
35



$\omega(\text{Ca}) = 40\%$

$m(\text{CaCO}_3) = 220 \cdot 5 = 1100 \text{ г}$

$m(\text{Ca}) = 0,4 \cdot 1100 \text{ г} = 440 \text{ г}$ Ответ: $m(\text{Ca}) = 440 \text{ г}$



1,55 *my*

2 4,5 *for*

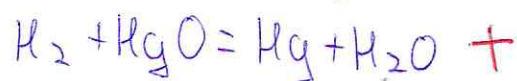
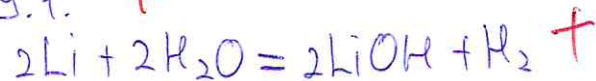
3-3,5 *near*

4 - 4 *d*

5 x *for*

1 55 м
2 45 м
3 15 м
4 — м
5 18 м
ШИФР 9-84

9.1.



9.2. Дано:

$m(\text{CaCO}_3)$ - в 1-ой банке 5 грамм

За 100 г 220 г

Найти:

$m(\text{Ca})$

Решение

$$\text{За } 200 \text{ г } 220 \cdot 5 = 1100 \text{ г. } \text{CaCO}_3$$

$$M(\text{CaCO}_3) = 40 + 12 + 48 = 100 \text{ г/моль}$$

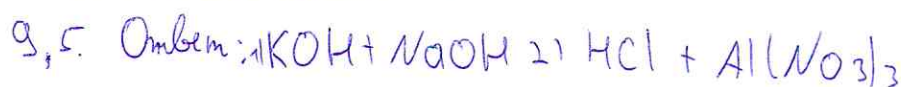
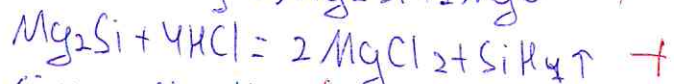
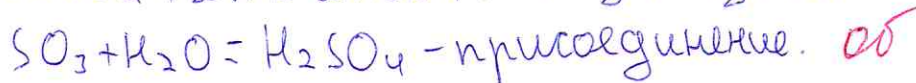
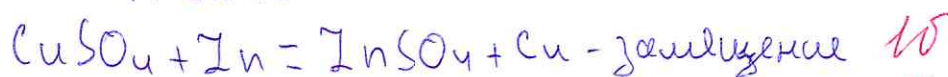
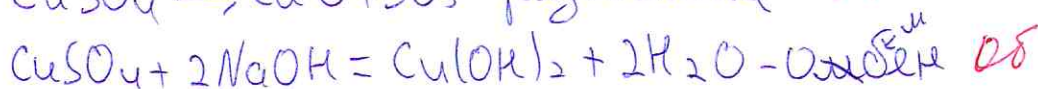
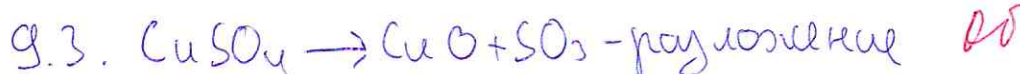
$$\text{в } 100 \text{ г. } \text{CaCO}_3 - 40 \text{ г. Ca}$$

$$\text{в } 1100 \text{ г. } \text{CaCO}_3 - x \text{ г. Ca}$$

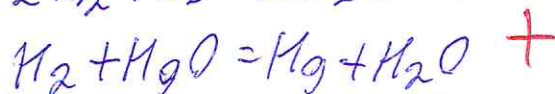
$$x = \frac{1100 \cdot 40}{100} = 11 \cdot 40 = 440 \text{ г.}$$

Ответ: 440 г.

45



10 Нет гр не и /



9.2. $m(\text{CaCO}_3) = 520 \text{ г}$
220 г и у за 20 г.
х г - ?

Решение:
 $M(\text{CaCO}_3) = 40 + 12 + 48 = 100$
За 20 г $220 : 5 = 1100 \text{ г. CaCO}_3$
В 100 г $\text{CaCO}_3 - 40 \text{ г. Ca}$
В 1100 г $\text{CaCO}_3 - x \text{ г. Ca}$
 $x = 11 \cdot 40 = 440 \text{ г.}$

Ответ: 440 г.

9.3. разложение:



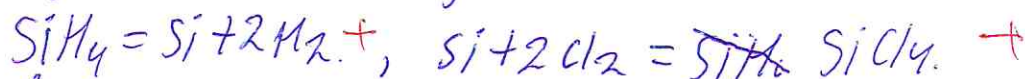
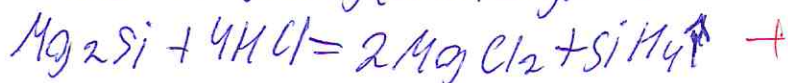
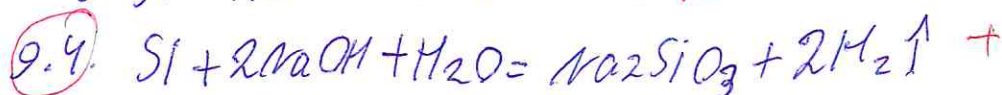
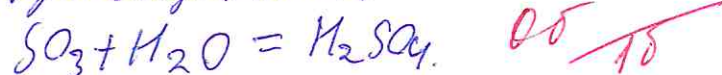
Однук:



замещение:



присоединение:



9.5. Можно только 2 пары.

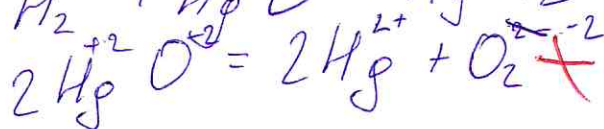
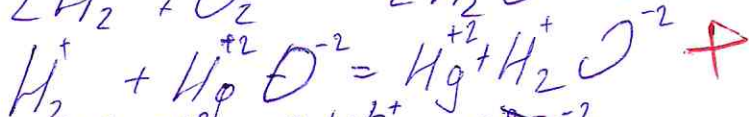
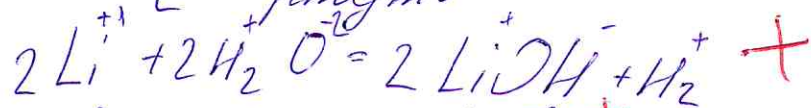


Нет уравнений 15

9.1 А - шитий

В - водород

С - ртуть



9.2

$$M(\text{CaCO}_3) = 100 \text{ г/моль}$$

$$w(\text{Ca}) = 40\%$$

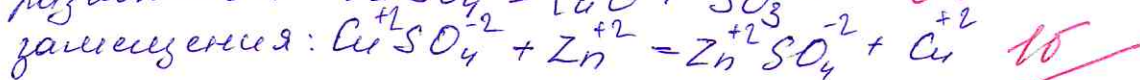
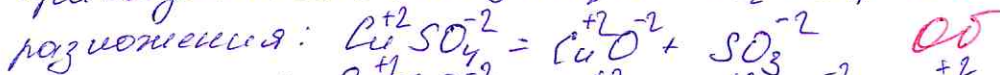
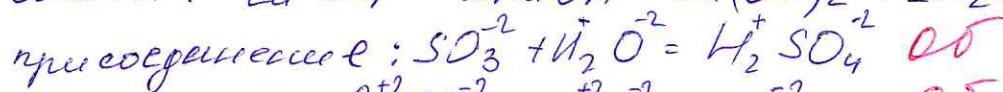
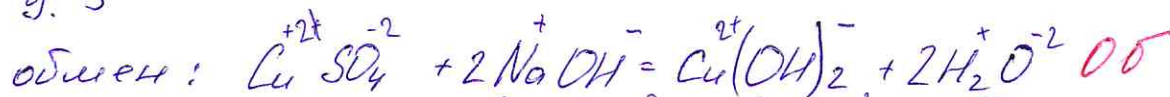
$$M(\text{CaCO}_3) = 220 \cdot 0,5 = 110 \text{ г}$$

$$M(\text{Ca}) = 1100 \cdot 0,4 = 440 \text{ г} - \text{необход. кол-во Ca}$$

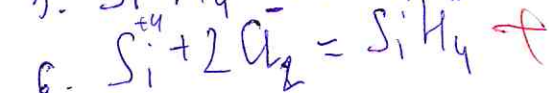
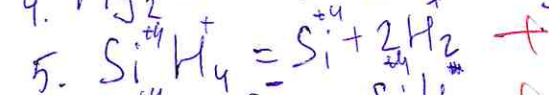
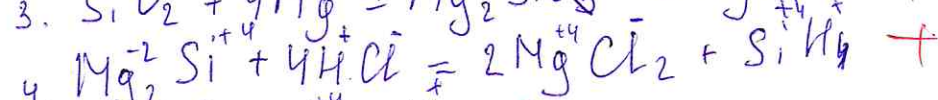
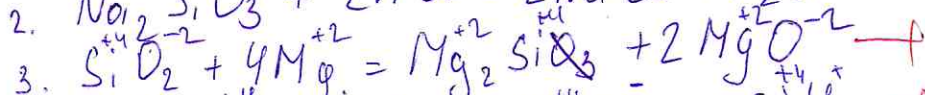
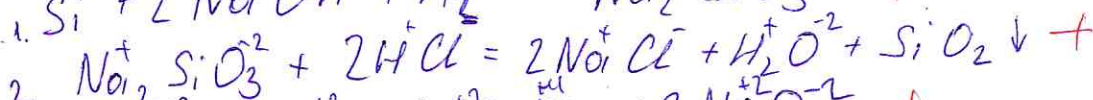
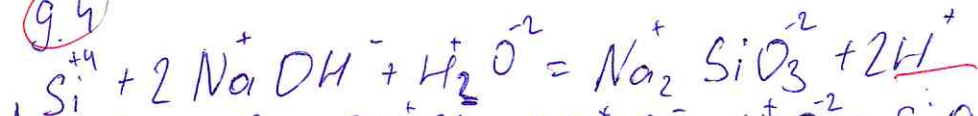
Ответ: 440 г

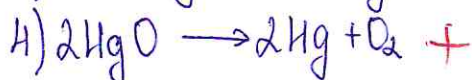
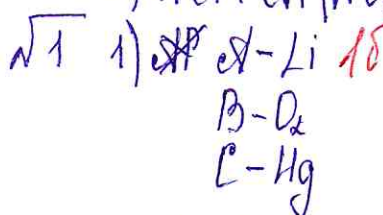
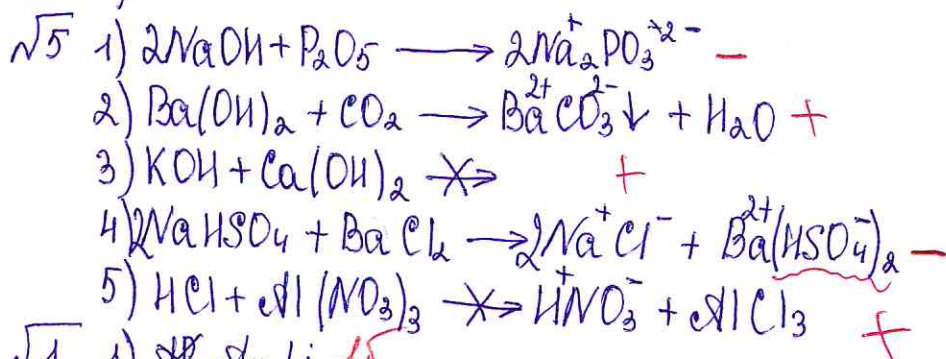
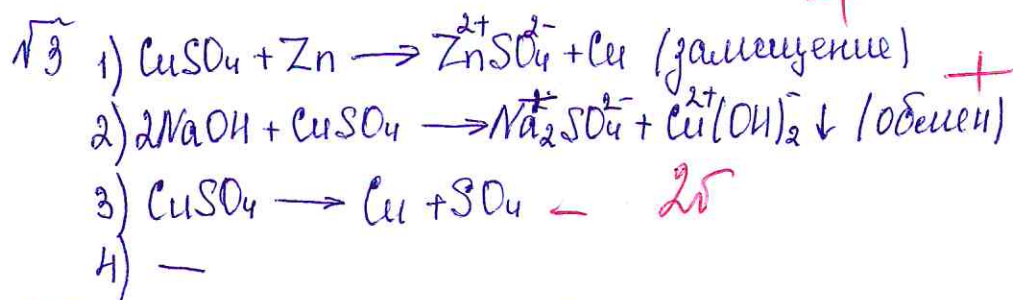
45

9.3



9.4





$\sqrt{2}$ Дано:

$\frac{m(\text{CaCO}_3)}{m(\text{Ca})} = 57$
 $m(\text{Ca}) = ?$

Решение:

$\omega = \frac{m_{\text{в.р.}}}{m} \cdot 100\%$

$m(\text{CaCO}_3) = 40 + 12 + 48 = 100 \text{ г/моль}$

$\nu(\text{CaCO}_3) = \frac{5}{100} = 0,05 \text{ моль}$

$\nu(\text{Ca}) : \nu(\text{CaCO}_3) = 1:1 \Rightarrow \nu(\text{Ca}) = \nu(\text{CaCO}_3)$

$m(\text{Ca}) = \nu(\text{Ca}) \cdot m(\text{Ca})$

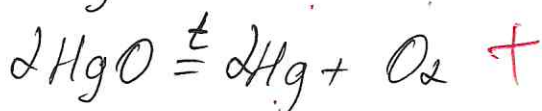
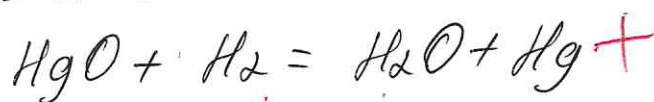
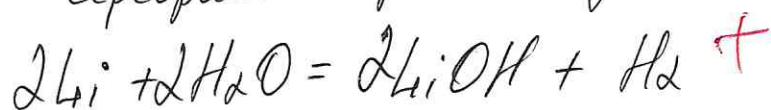
$m(\text{Ca}) = 0,05 \cdot 40 = 2 \text{ г} \Rightarrow 220 \cdot 2 = 440 \text{ г}$

Ответ: $m(\text{Ca}) = 440 \text{ г}$ 45

155 My D
 245 Sl D
 3-25 D M
 4 - x d M
 538 D - D -

36

9.1. Металл А - это Li (Литий) 15
 Газ В - это H₂ (Водород)
 Серебристо-серая жидкость С - это Hg (Ртуть)



1 15 ~~15~~
 2 45 ~~45~~
 3 25 ~~25~~
 4 - x ~~45~~
 5 45 ~~45~~

9.2. Дано:

$$m(CaCO_3) = 52$$

в 1 а.и.е.

$$m_{\text{ка-во а.и.е.}} = 220 \text{ шт.}$$

$$m(Ca) = ?$$

Решение:

$$\omega = \frac{n \cdot Ar(Ca)}{Mr} \cdot 100\%$$

$$Mr(CaCO_3) = 40 + 12 + 16 \cdot 3 = 100$$

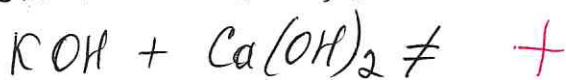
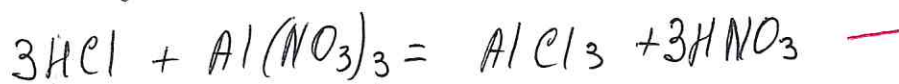
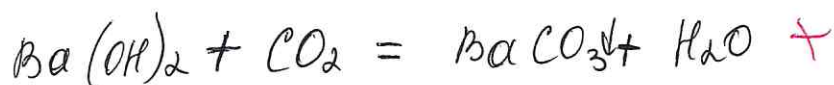
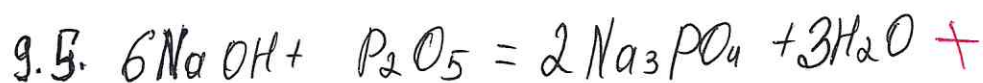
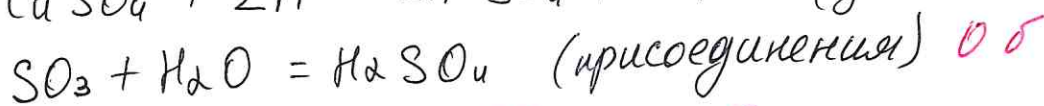
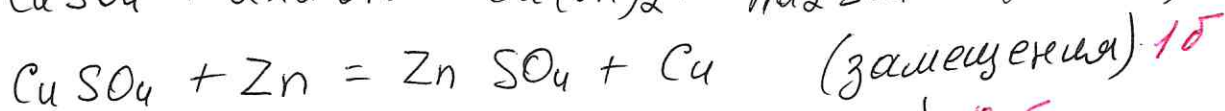
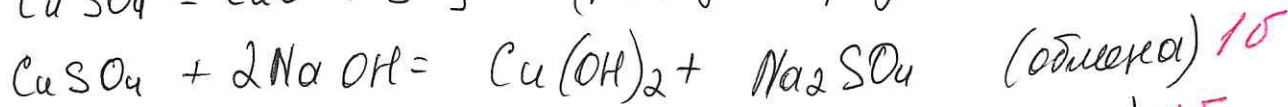
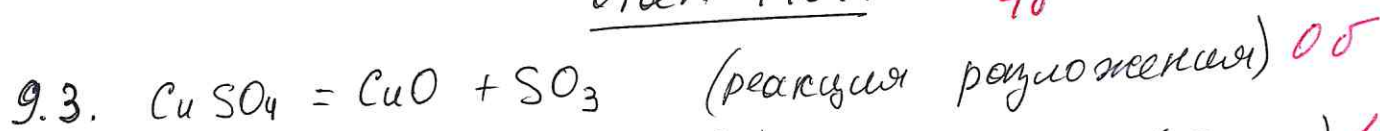
$$\omega = \frac{40}{100} \cdot 100\% = 40\%$$

$$m(Ca) \text{ в 1 а.и.е.} = 52 \cdot 40\% = 22$$

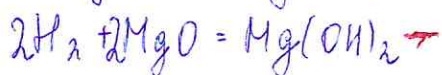
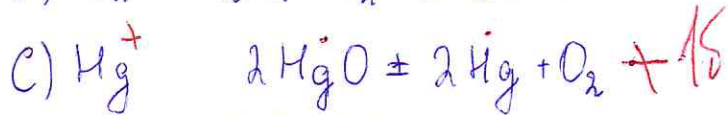
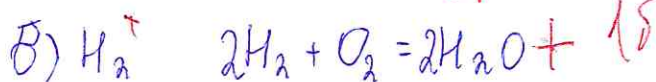
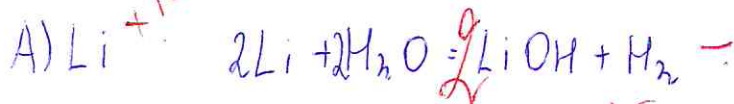
$$m(Ca) \text{ в 220 а.и.е.} = 2 \cdot 220 = 440$$

Ответ: 440г.

45



45

N1 ⁺¹⁵

1-35 *my*
 2-45 *ok*
 3-15 *max*
 4-2 *ok*
 5-45 *ok* -

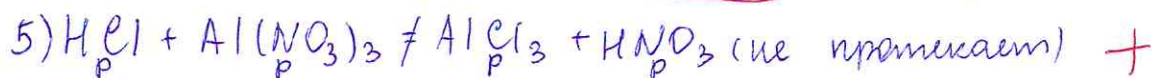
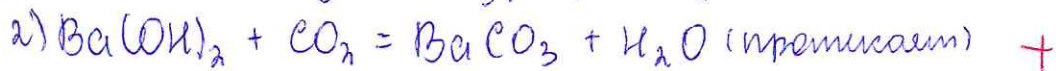
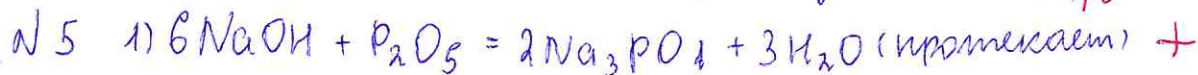
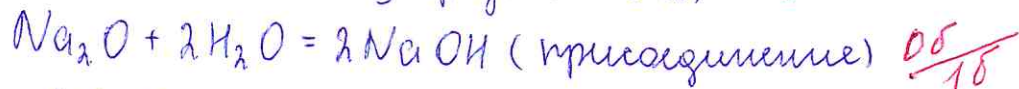
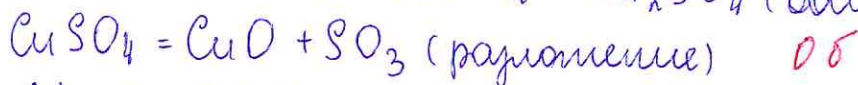
N2 $m(Ca) = w(Ca) \cdot m(CaCO_3)$

$w(Ca) = \frac{40}{100} = 0,4 \text{ масс}$

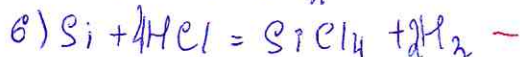
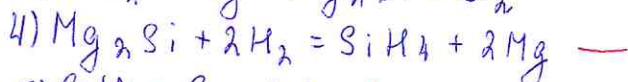
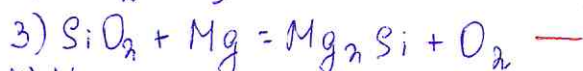
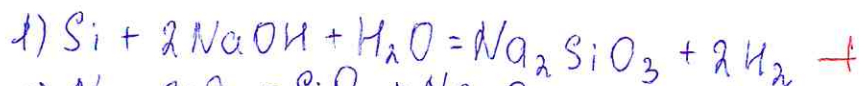
$m(Ca) = 0,4 \text{ масс} \cdot 52 = 22$

$m(Ca \text{ за год}) = 22 \cdot 20 = 4402$

Ответ: 4402 килограмм должно воспламениться в
 опашные пещники за год. 45

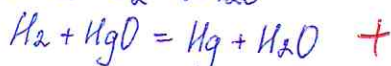
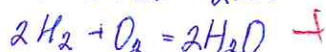
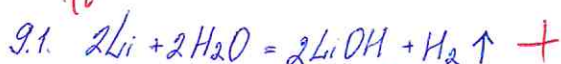


N4



15

Бланк ответа

Класс 9B.ШИФР 9-20

9.2. Дано:

$m(CaCO_3) = 5 \cdot 200$

$m(Ca) = ?$

$M(CaCO_3) = 40 +$

$12 + 16 \cdot 3 = 100$ г/моль

Решение:

$220 \text{ моль} - 12 \text{ моль}$

$1 \text{ моль} - 52 (CaCO_3)$

1) $220 \cdot 5 = 1100 (CaCO_3)$ потеря

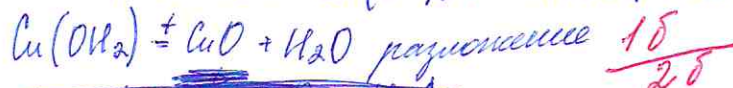
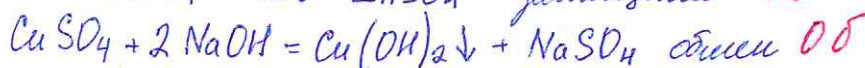
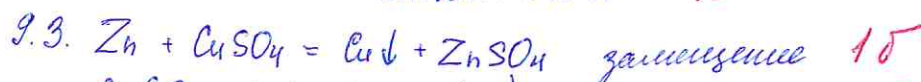
2) $w(Ca) = \frac{n \cdot A_r(Ca)}{M_r(CaCO_3)} = \frac{40}{100} = 0,4 = 40\%$

$1100 - 100\% CaCO_3$

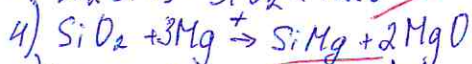
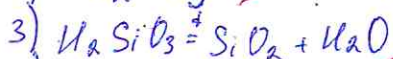
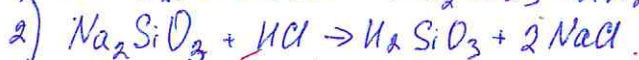
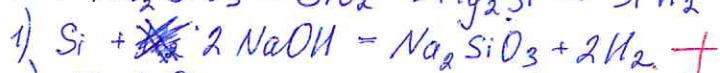
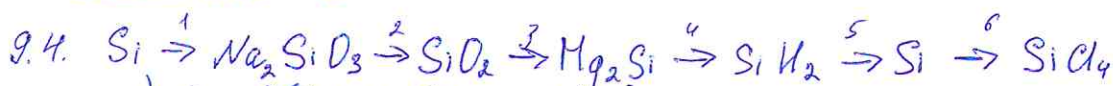
$x - 40\% Ca$

$x = \frac{1100 \cdot 40}{100} = 440$ г

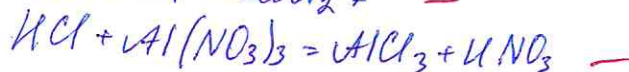
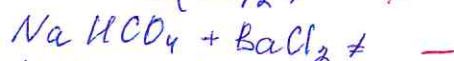
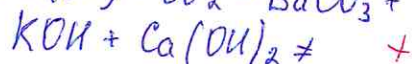
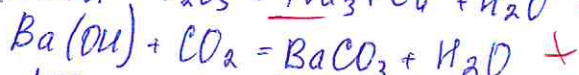
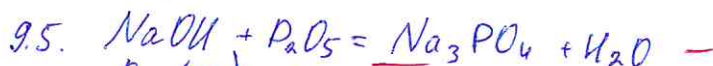
Ответ: 440 г. 45



25

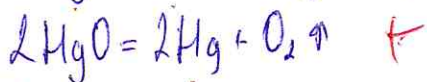
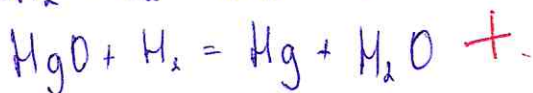
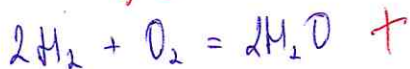
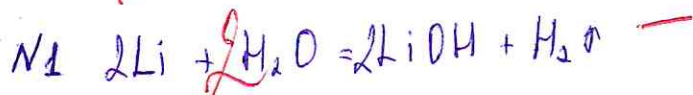


155



25

1 45 м/с
2 45 м/с
3 25 м/с
4 15 м/с
5 25 м/с



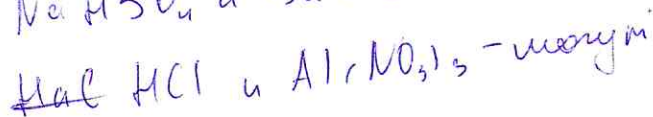
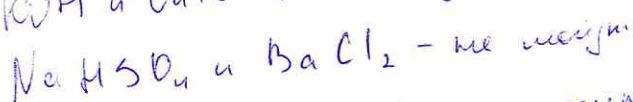
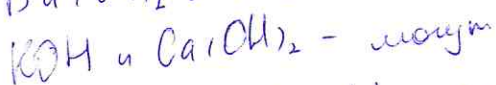
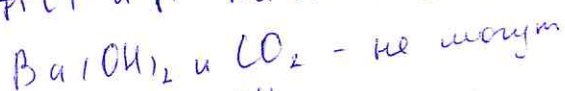
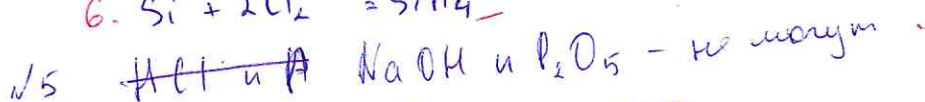
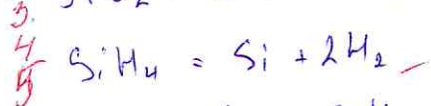
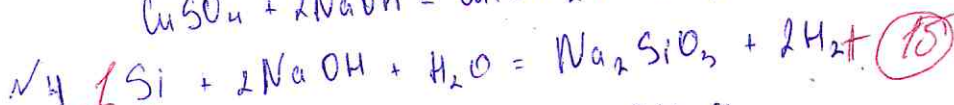
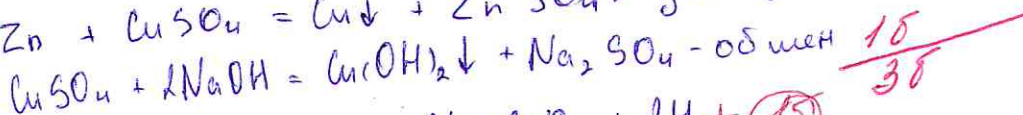
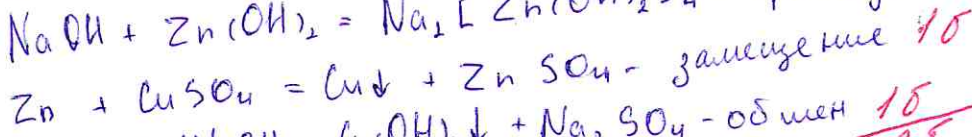
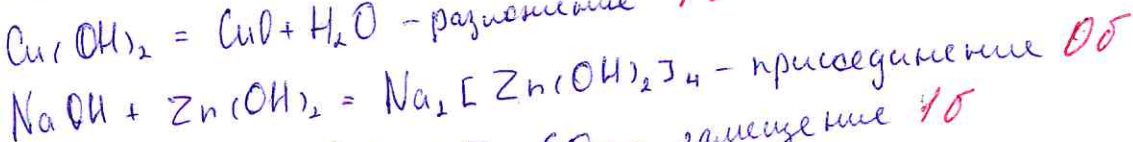
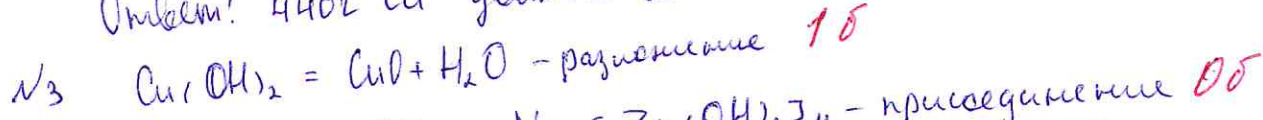
A — Li (литий), B — H₂ (водород), C — Hg — ртуть

N2. $w(\text{Ca}) = \frac{40}{100} = 0,4$

$m(\text{Ca}) = w(\text{Ca}) \cdot m(\text{CaCO}_3) = 0,4 \cdot 52 = 22 \text{ г}$

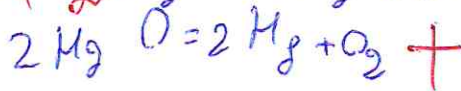
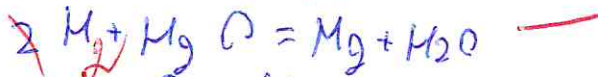
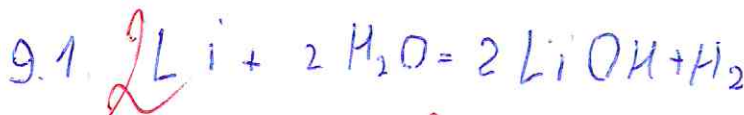
$22 \cdot 220 = 4840 \text{ г Ca}$ 35

Ответ: 4840 г Ca должна восстановится в органические соединения за 29



неб р-ции 25

1 15 M D
2 35-кар-не з
3 35 кар D
4 15 D
5 25 D



A - литий 15

B - водород

C - оксид ртути и ртути

1 - 25 шт. 15

2 - 45 шт. 15

3 - 15 шт. 15

4 - 65 шт. 15

5 - 0 шт. 15

15

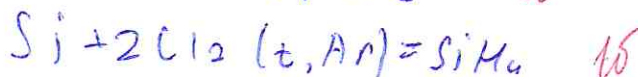
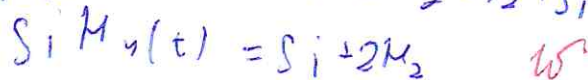
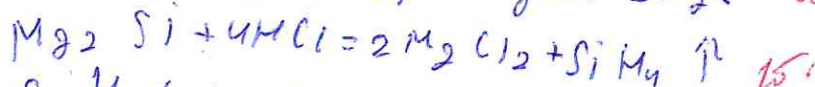
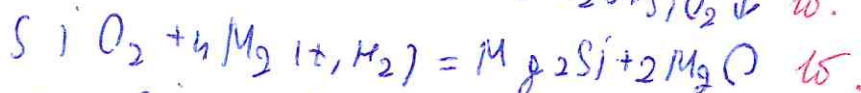
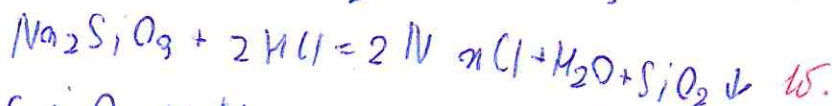
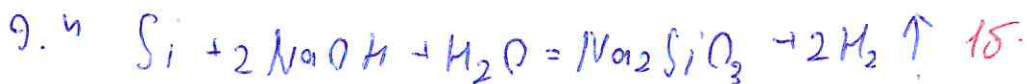
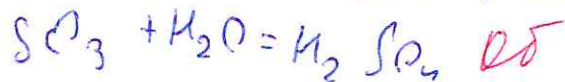
9.2.

$$M(CaCO_3) = 40 + 12 + 48 = 100$$

$$\frac{1002 - 402 Ca}{152 - x} \quad x = 5 \cdot \frac{40}{100} = 2$$

$$2 \cdot 220 = 440 \text{ г } Ca \quad 45$$

9.3.



(65)

9.5

а) 1-мет —

б) 1-мет —

в) 1-мет —

г) 1-мет —

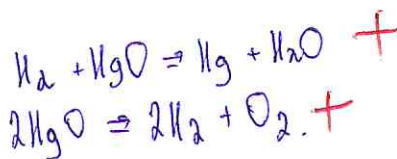
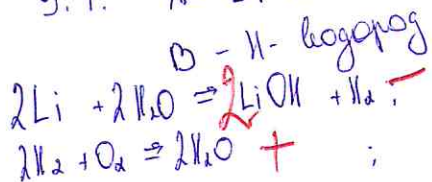
д) 1-мет —

е) 1-мет —

все уравнения 05

9.1. А - Li литий

С - Hg - ртуть.

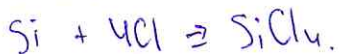
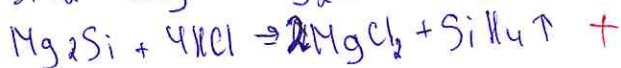
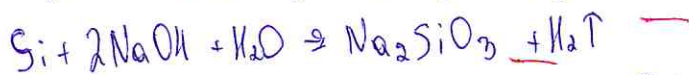
9.2. CaCO₃ = 40 + 12 + 48 = 100.Ca = 40; C = 12; O₃ = 16 · 3 = 48.в 52-60-й, процент; $x = (5 \cdot 40) : 100 = 2$.

2. 220 г. = 440 г.; 440 г. необходимо восстановить за 10 г.

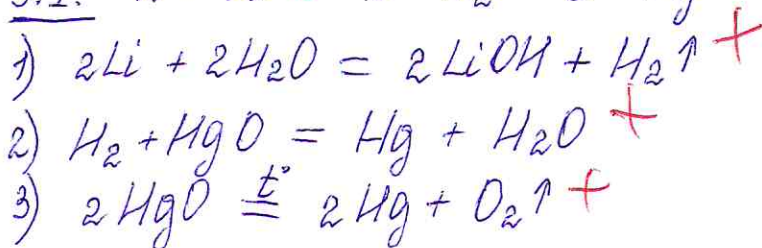
Ответ: 440 г.

9.3. NaOH; CuSO₄; Zn.Разложение: CuSO₄ ⇒ CuO + SO₃Обмен: CuSO₄ + 2NaOH ⇒ Cu(OH)₂ + 2H₂O.Замещение: CuSO₄ + Zn ⇒ ZnSO₄ + CuГидратация: SO₃ + H₂O ⇒ H₂SO₄.

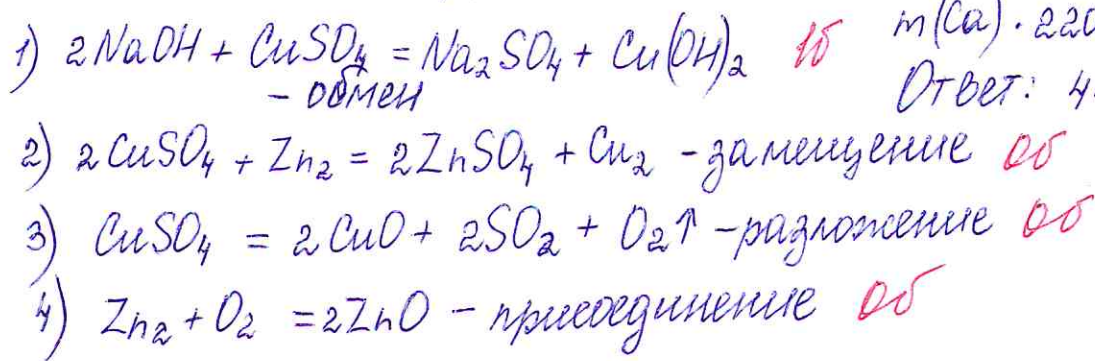
9.4



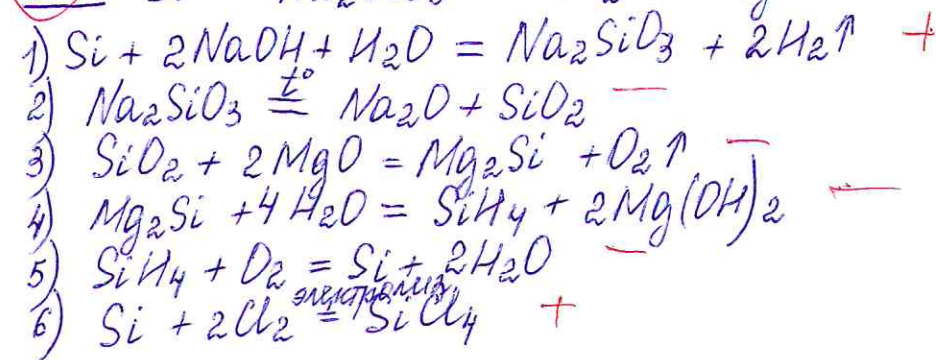
9.1. A - Li B - H₂ C - Hg



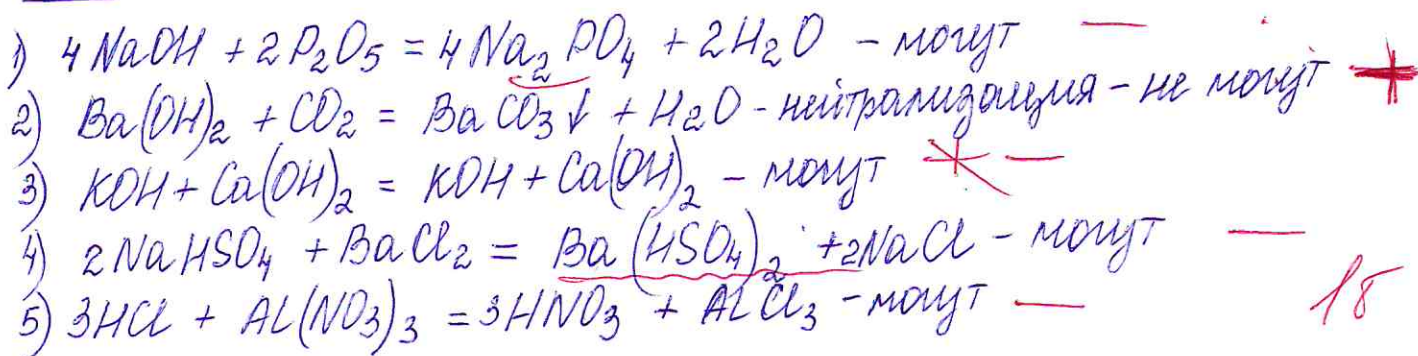
9.3. NaOH, CuSO₄, Zn₂



9.4. $\text{Si} \xrightarrow{1} \text{Na}_2\text{SiO}_3 \xrightarrow{2} \text{SiO}_2 \xrightarrow{3} \text{Mg}_2\text{Si} \xrightarrow{4} \text{SiH}_4 \xrightarrow{5} \text{Si} \xrightarrow{6} \text{SiCl}_4$



9.5.



1 40 48 100
 2 40 12 48
 3 100 40 12 48
 4 - 20 100
 5 18 100

9.2. Дано:

$$m(\text{CaCO}_3) = 52$$

$$m(\text{Ca}) = ?$$

за 100 - 220 яич

$$m(\text{Ca}) \cdot 220 = ?$$

$$+ A_r(\text{C}) + 3A_r(\text{O}) = 40 + 12 + 48 = 100$$

$$m(\text{Ca}) = 5 \cdot 0,4 = 2(2) - \text{на яиче}$$

$$m(\text{Ca}) \cdot 220 = 440(2) - \text{за 100}$$

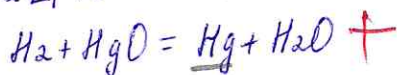
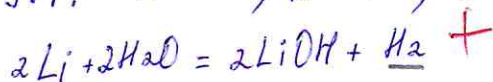
Ответ: 440 г.

Решение:

$$\omega(\text{Ca}) = \frac{A_r(\text{Ca})}{M_r(\text{CaCO}_3)} \cdot 100\%$$

$$= \frac{40}{100} = 0,4$$

$$M_r(\text{CaCO}_3) = A_r(\text{C})$$

9.1. А - Li, В - H₂, С - Hg.

9.2. Дано:

за 100 - 220 атом.

 $m(\text{CaCO}_3)$ в атом = 52. $m(\text{Ca})$ за 100 - ? г

Решение:

$$M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 52 + 48 = 100 \text{ г.}$$

$$M(\text{Ca}) = 40 \text{ г.}$$

$$100 \text{ г} = 40 \text{ г.}$$

$$52 = x \text{ г.}$$

$$x = \frac{40 \cdot 52}{100}$$

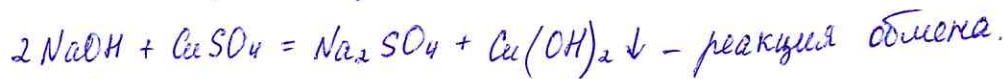
$$x = \frac{2080}{100}$$

$$x = 20,8 \text{ г. Ca в } 52 \text{ г. CaCO}_3.$$

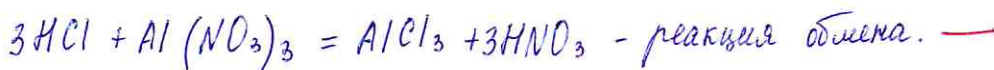
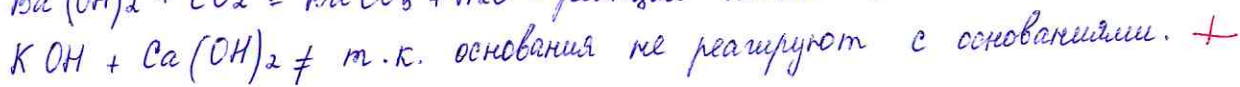
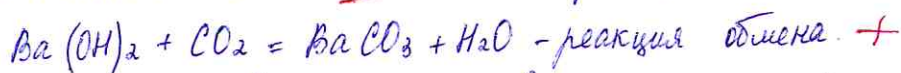
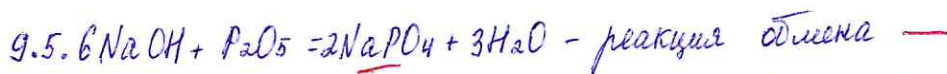
$$m(\text{Ca}) \text{ за } 100 = 220 \cdot 2 = 440 \text{ г.}$$

Ответ: 440 г.

45

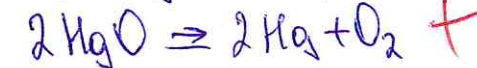
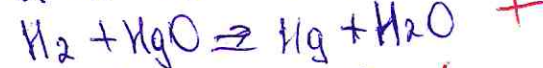
9.3. NaOH, CuSO₄, Zn.

25



35

1 35
2 45
3 25
4 -
5 35

9.1. А-Li В-Н₂ С-Нg

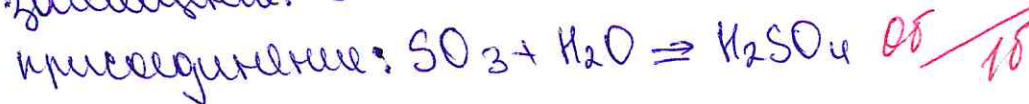
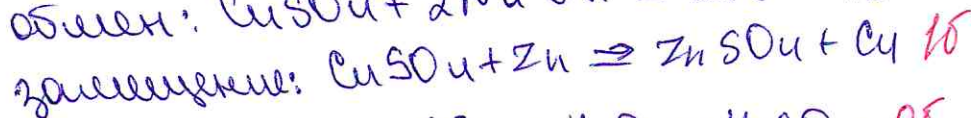
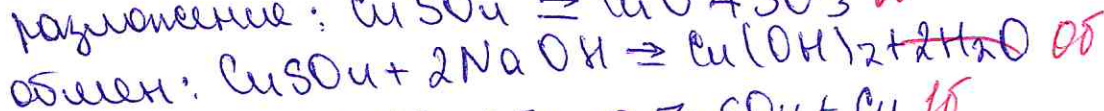
9.2.

$$m(\text{CaCO}_3) = 100 \text{ г/моль}$$

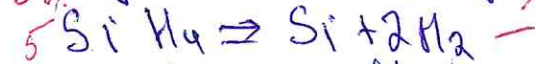
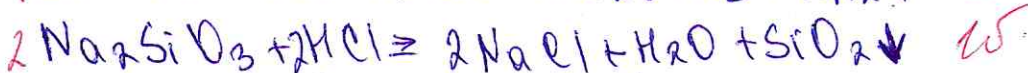
$$m(\text{CaCO}_3) = 220 \cdot 5 = 1100 \text{ г}$$

$$m(\text{Ca}) = 1100 \cdot 0,4 = 440 \text{ г} \quad 45$$

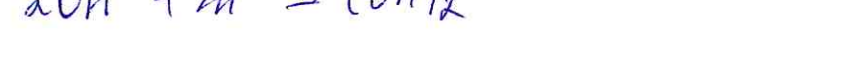
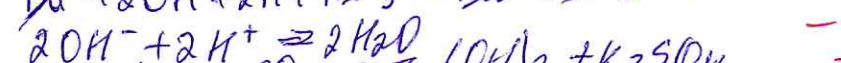
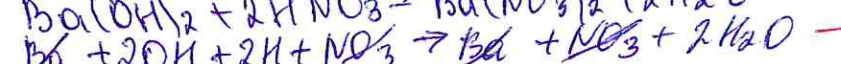
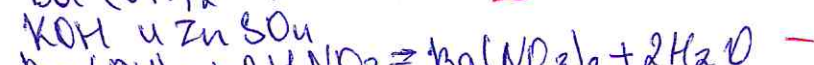
9.3.



9.4.



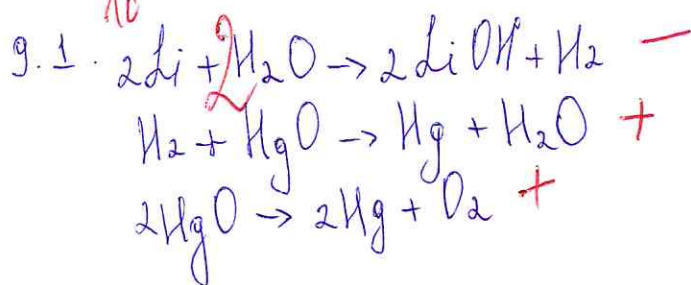
9.5.



155 М. В.
2 45
3 15 мар
4 - 25
5 08

(20)

08



1 - 35 М
 2 - 45
 3 - 25
 4 - 35
 5 - X

9.2. Дано:

$m(\text{CaCO}_3) = 5\text{г}$
 в одном литре
 кон-во суш =
 = 220

$m(\text{Ca}) = ?$

Решение:

За 100 г масса $\text{CaCO}_3 = 220 \cdot 5 = 1100\text{г}$.

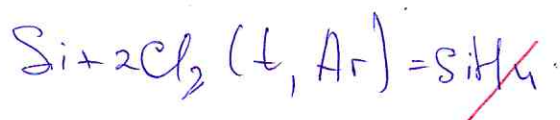
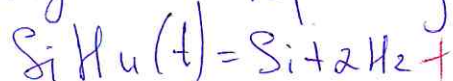
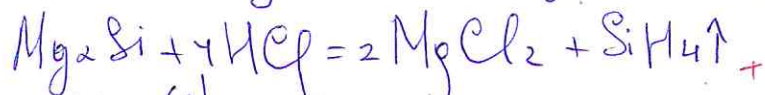
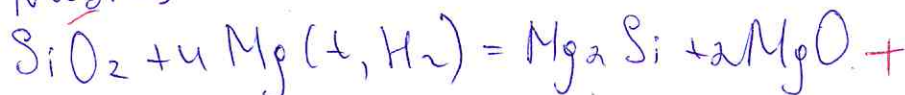
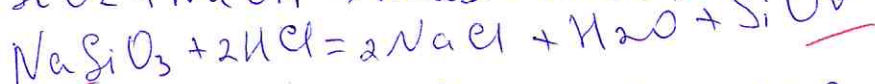
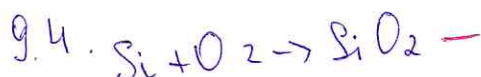
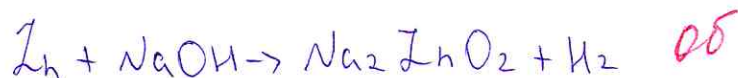
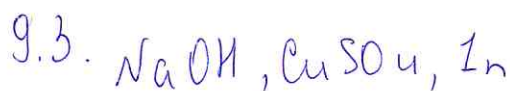
$M(\text{CaCO}_3) = 40 + 12 + 16 \cdot 3 = 100$

В 1 моле $\text{CaCO}_3 = 40\text{г Ca}$

$\nu(\text{CaCO}_3) = \frac{m}{M} = \frac{1100 \cdot 40}{100} = 440\text{ моле}$

Ответ: 440

45



35